



ACIDS, BASES & SALTS

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SCIENCE PRACTICE PAPER 02

Total Marks : 20

I. Choose the Most Appropriate Answers

3 x 1 = 3

- How many water molecules does hydrated calcium sulphate contain?
a. 5 b. 10 c. 7 d. 2
- An aqueous solution turns red litmus solution blue. Excess addition of which of the following solution would reverse the change?
a. Baking powder c. Ammonium hydroxide solution
b. Lime d. Hydrochloric acid
- In terms of acidic strength, which one of the following is in the correct increasing order?
a. Water < Acetic acid < Hydrochloric acid
b. Water < Hydrochloric acid < Acetic acid
c. Acetic acid < Water < Hydrochloric acid
d. Hydrochloric acid < Water < Acetic acid

II. Answer the following questions

2 x 1 = 2

- Name two substances which can be used as olfactory indicator
- Write the chemical formula of hydrated copper sulphate and anhydrous copper sulphate.

III. Answer the Following questions

4 x 2 = 8

- "Sodium hydrogencarbonate is a basic salt". Justify the statement. How is it converted into washing soda? Explain
- Explain why is hydrochloric acid a strong acid and acetic acid, a weak acid. How can it be verified?
- In the following list of acids, separate strong acids from weak acids. Hydrochloric acid, citric acid, acetic acid, nitric acid, formic acid, sulphuric acid

IV. Answer the following questions

x 3 = 3

- Describe an activity with diagram to illustrate that the reaction of metal carbonates and metal bicarbonates with acids produces carbon dioxide. Write the relevant equations of all the reactions that take place.

V. Answer the following questions

1 x 4 = 4

- A metal compound 'X' reacts with dil. H_2SO_4 to produce effervescence, The gas evolved extinguishes a burning candle. If one of the compound formed is calcium sulphate, then what is 'X' and the gas evolved? Also, write a balanced chemical equation for the reaction which occurred?